

(Group 7 – Halogens)

Q1. (2018)

7

Q2. (2018)

F₂

Q3. (2018)

the reactivity decreases (going
down Group 7)

(because) chlorine displaces
bromine and iodine

(and) bromine displaces iodine
or iodine does not react

Q4.(2019)

(formula) At₂

(state) solid

Q5. a) (2020)

gas

Q5. b) (2020)

-35 (°C)

Q5. c) (2020)

increase

increase

Q6.a) (2020)

chlorine gas is toxic

Q6.b) (2020)

increased

chlorine (atoms) are now part of
the solid (iron chloride)

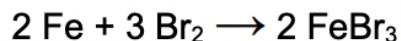
or

the mass of the chlorine (atoms)
is now also measured

Q6.c) (2020)

burns very vigorously

Q6.d) (2020)



Q6.e) (2020)

$$56 + (3 \times 80)$$

$$= 296$$

Q7.(2022)

Element	State at 150 °C	Symbol	Formula of the compound with hydrogen
Fluorine	gas	F	HF
Chlorine	gas	Cl	HCl
Bromine	gas	Br	HBr
Iodine	liquid	I	HI
Astatine	solid	At	HAt

Q8.(2022)



Q9.(2023)

(number of outer shell electrons)

7

(reason) (tennessine is in)

Group 7

Q10.(2024)

chlorine and potassium bromide

Q11.(2024)

relative molecular mass

increases and boiling point

increases