

(Covalent bonding and simple molecules)

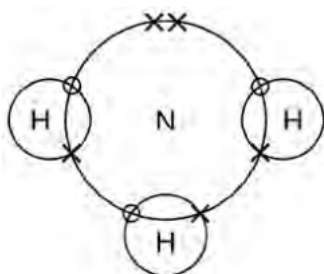
Q1. (2018)

hydrogen chloride is made of
small molecules

(so hydrogen chloride) has weak
intermolecular forces

(intermolecular forces) require
little energy to overcome

Q2. (2019)



Q3. (2019)

does not show the shape

or

only two-dimensional

Q4. (2019)

(ammonia has) small molecules

(ammonia has) weak
intermolecular forces

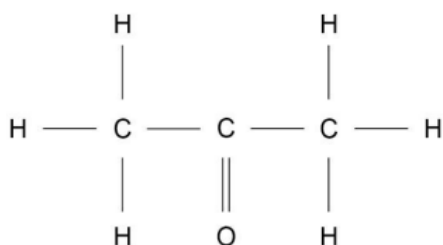
(so) little energy is needed to
overcome the intermolecular
forces

Q5. (2020)

poly(ethene)

water

Q6. a). (2021)



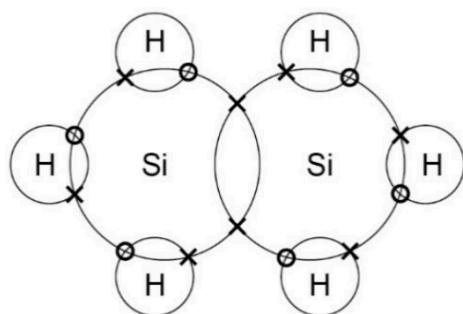
Q6. b). (2021)

$\text{C}_3\text{H}_6\text{O}$

Q7. (2021)

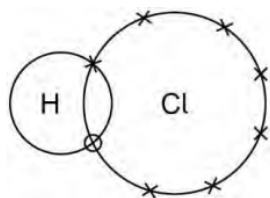
the intermolecular forces are
weak

Q8. (2022)



Q9. (2023)

an answer of



or

an answer of



Q10. (2023)

(methane)

methane has (much) smaller
molecules

(so) has weaker intermolecular
forces

(so the intermolecular forces)
need less energy to overcome

(so) the boiling / melting point is
lower (and methane is a gas)

Q11. (2024)

propane is a small molecule

(so) the forces between
molecules are weak

or

(so) the intermolecular forces
are weak

(which) require little energy to
overcome