

(Group 1 – Alkali metals)

Q1. (2018)

Iron is a transition metal.

Sodium is a Group 1 metal.

Give **two** differences between the properties of iron and sodium.

[2 marks]

1 _____

2 _____

Q2. (2019)

Sodium is in Group 1 of the modern periodic table.

Describe what you would see when sodium reacts with chlorine.

[2 marks]

Q3. (2021)

This question is about Group 1 elements.

Give **two** observations you could make when a small piece of potassium is added to water.

[2 marks]

1 _____

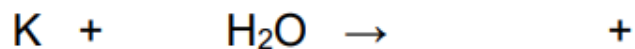
2 _____

Q4. (2021)

Complete the equation for the reaction of potassium with water.

You should balance the equation.

[2 marks]



Q5. (2021)

Explain why the reactivity of elements changes going down Group 1.

[4 marks]

Q6. (2022)

Give **two** differences between the physical properties of the elements in Group 1 and those of the transition elements.

[2 marks]

1

2

Q7. a). (2022)

This question is about groups in the periodic table.

The elements in Group 1 become more reactive going down the group.

Rubidium is below potassium in Group 1.

Rubidium and potassium are added to water.

Predict **one** observation you would see that shows that rubidium is more reactive than potassium.

[1 mark]

Q7. b). (2022)

Explain why rubidium is more reactive than potassium.

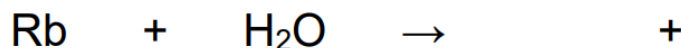
[3 marks]

Q7. c). (2022)

Complete the equation for the reaction of rubidium with water.

You should balance the equation.

[3 marks]



Q8. (2024)

This question is about the periodic table.

Sodium and potassium are in Group 1 of the periodic table.

Give **one** similarity and **one** difference between the electronic structures of sodium and potassium.

[2 marks]

Similarity _____

Difference _____

Q9. (2024)

Group 1 elements react with water.

Give **two** observations made when potassium reacts with water.

[2 marks]

1 _____

2 _____
