

(Relative atomic mass, isotopes, and abundance)

Q1. (2018)

isotope

Q2. (2018)

there are the same number of

$^{79}_{35}\text{Br}$ atoms and $^{81}_{35}\text{Br}$ atoms

Q3. (2020)

$$\frac{(203 \times 30) + (205 \times 70)}{100}$$

or

$$\frac{6090 + 14\,350}{100}$$

$$= 204.4$$

Q4. (2021)

**(atoms with the) same number
of protons**

**(but with) different numbers of
neutrons**

Q5. (2021)

$$(A_r =) \frac{(69 \times 60) + (71 \times 40)}{100}$$

$$= 69.8$$

Q6.a) (2022)

$$\begin{array}{l} \text{(relative atomic mass =)} \\ \frac{(39 \times 93.3) + (41 \times 6.7)}{100} \end{array}$$

$$= 39.134$$

$$= 39.1$$

Q6.b) (2022)

potassium / K

Q7. (2022)

neutrons

Q8. (2023)

protons

neutrons

Q9.(2023)

($A_r =$)

$$\frac{(6 \times 7.6) + (7 \times 92.4)}{100}$$

$$= 6.924$$

$$= 6.9$$

Q10.(2024)

atoms with the same number of
protons

with different numbers of
neutrons